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APPLICATION NO.	FILING DATE	FIRST NAMED INVENTOR	ATTORNEY DOCKET NO.	CONFIRMATION NO.
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10/593,772

09/22/2006

Go Ono

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BIRCH STEWART KOLASCH & BIRCH
PO BOX 747
FALLS CHURCH, VA 22040-0747

EXAMINER

CLARK, GREGORY D

ART UNIT

PAPER NUMBER

1786

NOTIFICATION DATE

DELIVERY MODE

11/10/2010

ELECTRONIC

Please find below and/or attached an Office communication concerning this application or proceeding.

The time period for reply, if any, is set in the attached communication.

Notice of the Office communication was sent electronically on above-indicated "Notification Date" to the following e-mail address(es):

mailroom@bskb.com

Office Action Summary	Application No. 10/593,772	Applicant(s) ONO ET AL.	
	Examiner GREGORY CLARK	Art Unit 1786	

-- The MAILING DATE of this communication appears on the cover sheet with the correspondence address --

Period for Reply

A SHORTENED STATUTORY PERIOD FOR REPLY IS SET TO EXPIRE 3 MONTH(S) OR THIRTY (30) DAYS, WHICHEVER IS LONGER, FROM THE MAILING DATE OF THIS COMMUNICATION.

- Extensions of time may be available under the provisions of 37 CFR 1.136(a). In no event, however, may a reply be timely filed after SIX (6) MONTHS from the mailing date of this communication.
- If NO period for reply is specified above, the maximum statutory period will apply and will expire SIX (6) MONTHS from the mailing date of this communication.
- Failure to reply within the set or extended period for reply will, by statute, cause the application to become ABANDONED (35 U.S.C. § 133). Any reply received by the Office later than three months after the mailing date of this communication, even if timely filed, may reduce any earned patent term adjustment. See 37 CFR 1.704(b).

Status

- 1) ☒ Responsive to communication(s) filed on 13 September 2010.
- 2a) ☐ This action is **FINAL**. 2b) ☒ This action is non-final.
- 3) ☐ Since this application is in condition for allowance except for formal matters, prosecution as to the merits is closed in accordance with the practice under *Ex parte Quayle*, 1935 C.D. 11, 453 O.G. 213.

Disposition of Claims

- 4) ☒ Claim(s) 1-11 is/are pending in the application.
- 4a) Of the above claim(s) _____ is/are withdrawn from consideration.
- 5) ☐ Claim(s) _____ is/are allowed.
- 6) ☒ Claim(s) 1-11 is/are rejected.
- 7) ☐ Claim(s) _____ is/are objected to.
- 8) ☐ Claim(s) _____ are subject to restriction and/or election requirement.

Application Papers

- 9) ☐ The specification is objected to by the Examiner.
- 10) ☐ The drawing(s) filed on _____ is/are: a) ☐ accepted or b) ☐ objected to by the Examiner.
Applicant may not request that any objection to the drawing(s) be held in abeyance. See 37 CFR 1.85(a).
Replacement drawing sheet(s) including the correction is required if the drawing(s) is objected to. See 37 CFR 1.121(d).
- 11) ☐ The oath or declaration is objected to by the Examiner. Note the attached Office Action or form PTO-152.

Priority under 35 U.S.C. § 119

- 12) ☐ Acknowledgment is made of a claim for foreign priority under 35 U.S.C. § 119(a)-(d) or (f).
- a) ☐ All b) ☐ Some * c) ☐ None of:
1. ☐ Certified copies of the priority documents have been received.
 2. ☐ Certified copies of the priority documents have been received in Application No. _____.
 3. ☐ Copies of the certified copies of the priority documents have been received in this National Stage application from the International Bureau (PCT Rule 17.2(a)).

* See the attached detailed Office action for a list of the certified copies not received.

Attachment(s)

- | | |
|--|---|
| 1) <input checked="" type="checkbox"/> Notice of References Cited (PTO-892) | 4) <input type="checkbox"/> Interview Summary (PTO-413) |
| 2) <input type="checkbox"/> Notice of Draftsperson's Patent Drawing Review (PTO-948) | Paper No(s)/Mail Date. _____ |
| 3) <input checked="" type="checkbox"/> Information Disclosure Statement(s) (PTO/SB/08) | 5) <input type="checkbox"/> Notice of Informal Patent Application |
| Paper No(s)/Mail Date <u>09/13/2010</u> . | 6) <input type="checkbox"/> Other: _____ |

DETAILED ACTION

Continued Examination Under 37 CFR 1.114

1. A request for continued examination under 37 CFR 1.114, including the fee set forth in 37 CFR 1.17(e), was filed in this application after final rejection. Since this application is eligible for continued examination under 37 CFR 1.114, and the fee set forth in 37 CFR 1.17(e) has been timely paid, the finality of the previous Office action has been withdrawn pursuant to 37 CFR 1.114. Applicant's submission filed on 09/13/2010 has been entered. Claims 1-12 pending.

Claim Rejections - 35 USC § 103

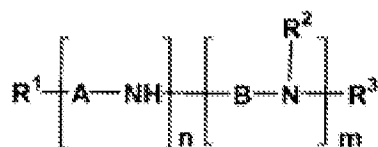
The following is a quotation of 35 U.S.C. 103(a) which forms the basis for all obviousness rejections set forth in this Office action:

(a) A patent may not be obtained though the invention is not identically disclosed or described as set forth in section 102 of this title, if the differences between the subject matter sought to be patented and the prior art are such that the subject matter as a whole would have been obvious at the time the invention was made to a person having ordinary skill in the art to which said subject matter pertains. Patentability shall not be negated by the manner in which the invention was made.

1. **Claims 1-9 are rejected under 35 U.S.C. 103(a) as being unpatentable over Oka (JP-07-090179) in view of Nalwa (Journal of Materials Science, 26 (1991) p. 1683-1690).**

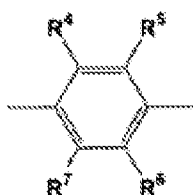
2. **Regarding Claims 1-3**, the applicant also claims a charge transporting varnish represented by Formula 1:

Art Unit: 1786

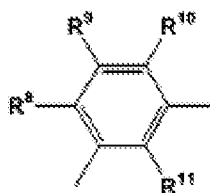


Formula 1

where R¹-R² are H, monovalent hydrocarbon group or organoxy group, R³ is an unsubstituted or substituted aryl group, A and B are divalent groups represented by Formula(s) 2 or 3:

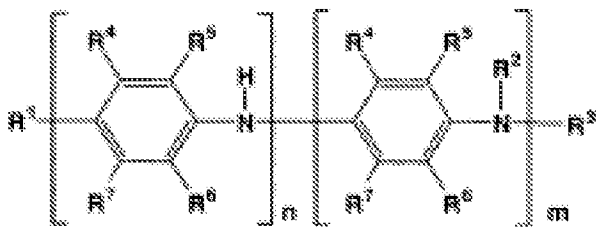


Formula 2



Formula 3

Applicants' formula 1 can also be represented by formula 4:



formula 4

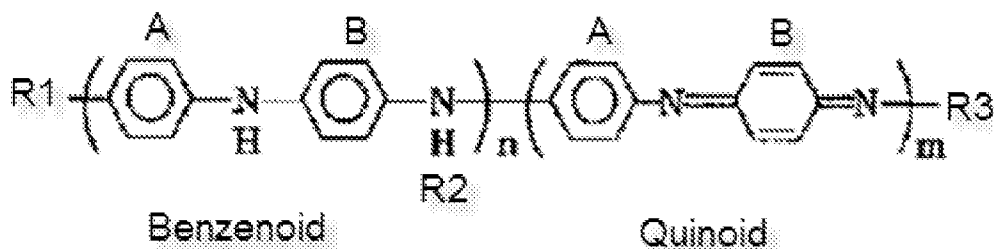
Wherein formula can also be represented as the quinonediimine derivative.

Oka discloses a charge transporting substance made of a charge transporting polyaniline and a polyimide (abstract).

Oka also discloses a charge transporting substance represented the polyaniline

Formula O-1 (below):

Art Unit: 1786



Formula O-1

Oka fails to mention a representative for R1 or the end-capping group, R3. The applicant claims R1 can be hydrogen and R3 as substituted or unsubstituted aryl.

Nalwa discloses that an oligoaniline skeleton contains both the benzenoid and quinoid (referred as quinonediimine by applicant) ring structures (abstract) (as seen above in Oka's formula O-1). The applicants' only shows the benzenoid in formula 4.

Oka's formula O-1 shows A and B are phenyl groups which correspond to the phenyl groups in formula 4 (per claim 3).

Oka fails to mention the R1 group.

The examiner takes the position that the phenyl ring to the far left in formula O-1 is the terminal ring position and the ring hydrogen para to the NH moiety corresponds to applicants' R1 group.

Oka fails to mention phenyl as an end-capping group.

Nalwa discloses that an oligoanilines can be end-capped with a phenyl group (page 1689).

The examiner takes the position that the nitrogen on the far right in the quinoid portion of formula O-1 represents the reactive site for propagation.

Whereas Oka does not mention the means to terminate the oligomer growth and Nalwa discloses the phenyl group can be used as an end-capping group, it would have

Art Unit: 1786

been obvious to a person of ordinary skills in the art at the time of the invention to have selected from known end-capping groups which would have included the phenyl group as disclosed by Nalwa which reads on applicants' limitation for the R3 position, absent unexpected results (per claim 2).

Oka discloses that the polyaniline-polyimide (polyamide acid) is soluble in a solvent (paragraph 8), is doped (paragraph 3) and has an average molecular weight of 2000 to 500,000 (abstract). Oka also discloses that the molecular weight of the polyaniline-polyimide complex is kept in a range have suitable solubility in a solvent (paragraphs 6 and 7). The applicant claims an average molecular weight of 250 to 800. Oka fails to teach a range which overlap applicants' range.

The average molecular weight is ultimately controlled by the propagation level and the molecular weight is view as a cause effective variable for solubility of the charge transporting varnish. A lower molecular weight materials would be expected to have higher solubility.

As Oka does not teach away from using lower molecular weight materials and such materials would be expected to improve solubility, one could envisage a skilled artisan making a series of materials at various molecular weights to determine which range offered suitable solubility properties.

With a reasonable expectation of success, a person of ordinary skill in the art at the time of the invention would test a series of polymers with varying propagation levels to determine which range gave suitable solubility properties which would have included the claimed range, absent unexpected results.

3. **Regarding Claim 4**, Oka discloses that the repeat unit values for n and m are as follows: $m + n = 10-5000$. The applicant claims $m+n$ is less than or equal to 6.

Oka explains that molecular weight affects the ultimate solubility of the polymer. Polymers with molecular weights that are too high would be expected to have lower solubility in a solvent.

With a reasonable expectation of success, a person of ordinary skill in the art would test a series of polymers with varying (different value of $m + n$) molecular weight ranges to determine which range give suitable solubility properties which would have included the range claimed by the applicant, absent unexpected results.

4. **Regarding Claims 5**, Oka discloses a weight ratio of polyaniline to polyamide is 1:0.1 - 1:10 which equates to a polyaniline weight range of 90% (1/1.1) to 9% (1/11) (paragraph 12). The applicant claims a % wt. of 0.1 to 50 wt. %.

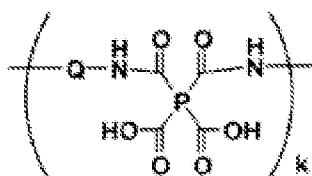
The range taught by Oka is not exactly the same as that of applicant however there is substantial overlap.

It would have been obvious to a person of ordinary skill in the art at the time of the invention to have adjusted to weight range to optimize the solubility of the oligomer which would have included the overlapping portion of the range, absent unexpected results.

Art Unit: 1786

5. **Regarding Claims 6**, Oka discloses that the polyimide precursor is represented by O-1 (shown below) with an average molecular weight of 1000-500,000. The applicant claims 1,000-50,000.

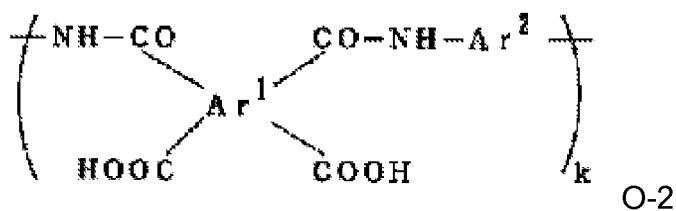
6. **Regarding Claim 7**, the applicant claims a polyimide precursor represented by Formula 8:



Formula 5

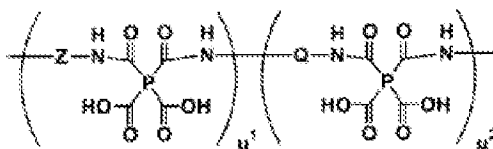
where Q can be a divalent phenyl group and P can be a tetravalent phenyl group.

Oka discloses formula O-2 shows a tetravalent organic group corresponding to applicants' P group and a divalent Ar² group corresponding to applicants' Q group (paragraph 4).



O-2

7. **Regarding Claims 8 and 9**, the applicant claims a polyimide precursor represented by formula 21:



Formula 21

Z and Q are divalent groups. Z is selected from applicants' formula(s) 22-27.

Oka discloses a polyimide precursor represented by O-2 (shown above). Oka also discloses that the polyimide precursor can be made by the reaction of a tetracarboxylic dianhydride or tetracarboxylic ester or tetracarboxylic acid halide with a diamine (paragraphs 9 and 10).

Oka further discloses that other diamines based on phenylenediamine and diaminobiphenyl (Z groups, the class of compounds listed in applicants' structures 22-27 in claim 9) can be used to make polyimides precursors that are analogous to applicants' Formula 21. The phenylenediamine or diaminobiphenyl compounds can have the following groups attached: methylenebis (aniline), bis (aminophenoxy) benzene, bis (aminophenoxy) biphenyl and bis [(aminophenoxy) phenyl] propane.

Oka also indicates the above groups can be further substituted (defined as W in the applicants' structures 22-27) with halogens, alkoxy (methoxy, ethoxyl, propoxy) groups, or alkyl groups (methyl, ethyl, propyl) (paragraph 9 and 10).

In essence, Oka discloses a polyimide precursor that can contain both Z and Q groups.

It would have obvious to a person of ordinary skill in the art at the time of the invention to have used a combination of diamine species defined as Z and Q groups in

Art Unit: 1786

Formula 21 to make suitable polyimide precursors through a reaction with a tetracarboxylic derivative.

Oka fails to disclose a propagation level of $u1/(u1 + u2)$ as greater than or equal to 0.2.

Oka discloses that the molecular weight is kept in a range to have suitable solubility in a solvent (paragraphs 6 and 7). The average molecular weight is ultimately controlled by the propagation level.

With a reasonable expectation of success, a person of ordinary skill in the art at the time of the invention would test a series of polymers with varying propagation levels to determine which range gave suitable solubility properties which would have included the propagation level claimed by the applicant, absent unexpected results.

8. Claims 10-11 are rejected under 35 U.S.C. 103(a) as being unpatentable over Oka (JP-07-090179) in view of Nalwa (Journal of Materials Science, 26 (1991) p. 1683-1690) and Kin (JP-11-185962).

9. Regarding Claims 10 and 11, Oka and Nalwa teach the invention of Claim 1 but fail to mention a charge transporting thin film or usage in an organic electroluminescent device.

It is common in the art to form a thin film layer from polyanilines for use in electroluminescent devices.

Art Unit: 1786

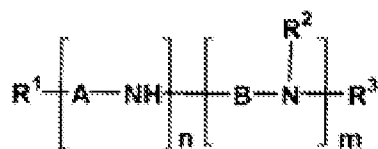
Kin discloses a doped polyaniline (charge transporting substance) (paragraphs 28-30) and the dopant can be benzenesulfonic acid (charge transporting dopant) (paragraph 30). The polymers are soluble in organic solvent (paragraph 39).

Kin also discloses a thin film layer (charge transporting varnish) (paragraph 32) as a component of an electroluminescent device (abstract).

10. As Kin discloses a similar type of conducting polyaniline material, it would have been obvious to a person of ordinary skill in the art at the time of the invention to have used the polyaniline materials of Oka as a thin film layer in an electroluminescent device, absent unexpected results.

11. **Claims 1-9 are rejected under 35 U.S.C. 103(a) as being unpatentable over Oka (JP-07-090179) in view of Nalwa (Journal of Materials Science, 26 (1991) p. 1683-1690) and Cameron (WO 90/11317).**

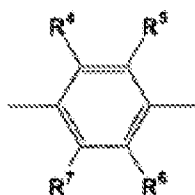
12. **Regarding Claims 1-3,** the applicant also claims a charge transporting varnish represented by Formula 1:



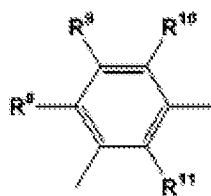
Formula 1

where R1-R2 are H , monovalent hydrocarbon group or organoxy group, R3 is an unsubstituted or substituted aryl group, A and B are divalent groups represented by Formula(s) 2 or 3:

Art Unit: 1786

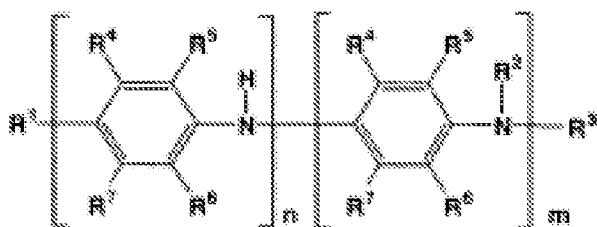


Formula 2



Formula 3

Applicants' formula 1 can also be represented by formula 4:



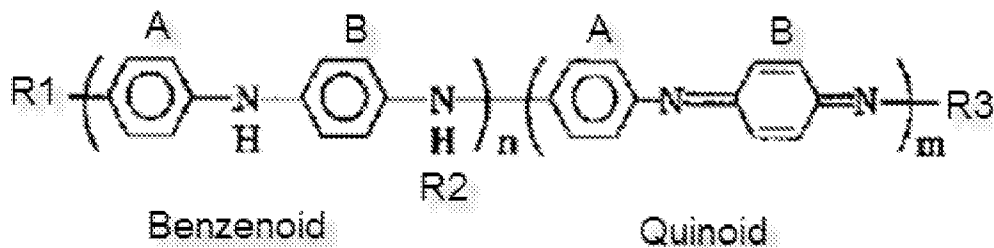
formula 4

Wherein formula can also be represented as the quinonediimine derivative.

Oka discloses a charge transporting substance made of a charge transporting polyaniline and a polyimide (abstract).

Oka also discloses a charge transporting substance represented the polyaniline

Formula O-1 (below):



Formula O-1

Oka fails to mention a representative for R1 or the end-capping group, R3. The applicant claims R1 can be hydrogen and R3 as substituted or unsubstituted aryl.

Nalwa discloses that an oligoaniline skeleton contains both the benzenoid and quinoid (referred as quinonediimine by applicant) ring structures (abstract) (as seen above in Oka's formula O-1). The applicants' only shows the benzenoid in formula 4.

Oka's formula O-1 shows A and B are phenyl groups which correspond to the phenyl groups in formula 4 (per claim 3).

Oka fails to mention the R1 group.

The examiner takes the position that the phenyl ring to the far left in formula O-1 is the terminal ring position and the ring hydrogen para to the NH moiety corresponds to applicants' R1 group.

Oka fails to mention phenyl as an end-capping group.

Nalwa discloses that an oligoanilines can be end-capped with a phenyl group (page 1689).

The examiner takes the position that the nitrogen on the far right in the quinoid portion of formula O-1 represents the reactive site for propagation.

Whereas Oka does not mention the means to terminate the oligomer growth and Nalwa discloses the phenyl group can be used as an end-capping group, it would have been obvious to a person of ordinary skills in the art at the time of the invention to have selected from known end-capping groups which would have included the phenyl group as disclosed by Nalwa which reads on applicants' limitation for the R3 position, absent unexpected results (per claim 2).

Oka discloses that the polyaniline-polyimide (polyamide acid) is soluble in a solvent (paragraph 8), is doped (paragraph 3) and has an average molecular weight of

Art Unit: 1786

2000 to 500,000 (abstract). Oka also discloses that the molecular weight of the polyaniline-polyimide complex is kept in a range have suitable solubility in a solvent (paragraphs 6 and 7). The applicant claims an average molecular weight of 250 to 800. Oka fails to teach a range which overlap applicants' range.

Cameron discloses conductive polymers which include Formula C-1 (page 1):



The examiner noted that Formula C-1 is closely related to Formula O-1 of Oka (discussed above).

Cameron further discloses that high molecular weight polymer based on Formula C-1 with MW in the range of 50,00 to 80,000 can be used. Likewise, oligomers based on Formula C-1 with MW of 800 to 900 can be used.

Oka/ Nalwa and Cameron teach aniline based polymeric materials which are conductive. Oka mentions mention that solubility of such material in solvents and Cameron mentions that high and low molecular weight aniline based polymeric materials can be used as conductive materials.

It would have been obvious to one of ordinary skill in the art at the time of the invention to have selected from known conductive materials in a suitable molecular weight range to ensure proper solubility in the reaction solvent which would have included low molecular weight oligoaniline based materials such as Formula C-1 of Cameron and to have selected the overlapping portion of the molecular weight ranges, absent unexpected results.

13. **Regarding Claim 4**, Oka discloses that the repeat unit values for n and m are as follows: $m + n = 10-5000$. The applicant claims $m+n$ is less than or equal to 6.

Oka explains that molecular weight affects the ultimate solubility of the polymer. Polymers with molecular weights that are too high would be expected to have lower solubility in a solvent.

Cameron teaches Formula C-1 (above) where $n = 0.5$. As Cameron teaches that low MW materials as effective conductive materials (as discussed above), a person of ordinary skill in the art would test a series of polymers with varying (different value of $m + n$) molecular weight ranges to determine which range give suitable solubility properties which would have included ranges taught which overlap the range claimed by the applicant, absent unexpected results.

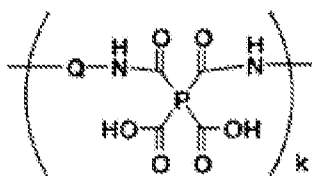
14. **Regarding Claims 5**, Oka discloses a weight ratio of polyaniline to polyamide is 1:0.1 - 1:10 which equates to a polyaniline weight range of 90% (1/1.1) to 9% (1/11) (paragraph 12). The applicant claims a % wt. of 0.1 to 50 wt. %.

The range taught by Oka is not exactly the same as that of applicant however there is substantial overlap.

It would have been obvious to a person of ordinary skill in the art at the time of the invention to have adjusted to weight range to optimize the solubility of the oligomer which would have included the overlapping portion of the range, absent unexpected results.

15. **Regarding Claims 6**, Oka discloses that the polyimide precursor is represented by O-1 (shown below) with an average molecular weight of 1000-500,000. The applicant claims 1,000-50,000.

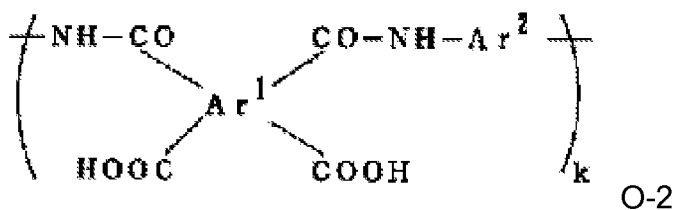
16. **Regarding Claim 7**, the applicant claims a polyimide precursor represented by Formula 8:



Formula 5

where Q can be a divalent phenyl group and P can be a tetravalent phenyl group.

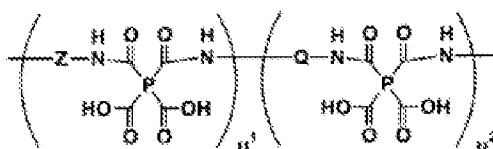
Oka discloses formula O-2 shows a tetravalent organic group corresponding to applicants' P group and a divalent Ar² group corresponding to applicants' Q group (paragraph 4).



O-2

Art Unit: 1786

17. **Regarding Claims 8 and 9**, the applicant claims a polyimide precursor represented by formula 21:



Formula 21

Z and Q are divalent groups. Z is selected from applicants' formula(s) 22-27.

Oka discloses a polyimide precursor represented by O-2 (shown above). Oka also discloses that the polyimide precursor can be made by the reaction of a tetracarboxylic dianhydride or tetracarboxylic ester or tetracarboxylic acid halide with a diamine (paragraphs 9 and 10).

Oka further discloses that other diamines based on phenylenediamine and diaminobiphenyl (Z groups, the class of compounds listed in applicants' structures 22-27 in claim 9) can be used to make polyimides precursors that are analogous to applicants' Formula 21. The phenylenediamine or diaminobiphenyl compounds can have the following groups attached: methylenebis (aniline), bis (aminophenoxy) benzene, bis (aminophenoxy) biphenyl and bis [(aminophenoxy) phenyl] propane.

Oka also indicates the above groups can be further substituted (defined as W in the applicants' structures 22-27) with halogens, alkoxy (methoxy, ethoxy, propoxy) groups, or alkyl groups (methyl, ethyl, propyl) (paragraph 9 and 10).

In essence, Oka discloses a polyimide precursor that can contain both Z and Q groups.

Art Unit: 1786

It would have obvious to a person of ordinary skill in the art at the time of the invention to have used a combination of diamine species defined as Z and Q groups in Formula 21 to make suitable polyimide precursors through a reaction with a tetracarboxylic derivative.

Oka fails to disclose a propagation level of $u1/(u1 + u2)$ as greater than or equal to 0.2.

Oka discloses that the molecular weight is kept in a range to have suitable solubility in a solvent (paragraphs 6 and 7) and Cameron teaches MW that overlap applicants' range as discussed above. The average molecular weight is ultimately controlled by the propagation level.

With a reasonable expectation of success, a person of ordinary skill in the art at the time of the invention would test a series of polymers with varying propagation levels to determine which range gave suitable solubility properties which would have included the propagation level claimed by the applicant, absent unexpected results.

18. Claims 10-11 are rejected under 35 U.S.C. 103(a) as being unpatentable over Oka (JP-07-090179) in view of Nalwa (Journal of Materials Science, 26 (1991) p. 1683-1690) and further in view of Cameron (WO 90/11317) and Kin (JP-11-185962).

Art Unit: 1786

19. **Regarding Claims 10 and 11**, Oka, Nalwa and Cameron teach the invention of Claim 1 but fail to mention a charge transporting thin film or usage in an organic electroluminescent device.

It is common in the art to form a thin film layer from polyanilines for use in electroluminescent devices.

Kin discloses a doped polyaniline (charge transporting substance) (paragraphs 28-30) and the dopant can be benzenesulfonic acid (charge transporting dopant) (paragraph 30). The polymers are soluble in organic solvent (paragraph 39).

Kin also discloses a thin film layer (charge transporting varnish) (paragraph 32) as a component of an electroluminescent device (abstract).

20. As Kin discloses a similar type of conducting polyaniline material, it would have been obvious to a person of ordinary skill in the art at the time of the invention to have used the polyaniline materials of Oka as a thin film layer in an electroluminescent device, absent unexpected results.

Response to Amendments

Applicant argues that there is no teaching in Oka for using a material with a molecular weight of below 2,000 and that Oka teaches away from lower MW materials.

The examiner counters that Oka teaches in paragraph 7 that MWs for the polyaniline polyimide complex below 2,000 eventually will lead to a loss in flexibility and MWs above 50,000 show lower solubility. The examiner interprets that to mean that the flexibility decreases as you drop below 2,000; however, there is no exact point

Art Unit: 1786

mentioned where said complex would not function as a conductive layer. Additionally, Cameron clearly teaches that high and low MW oligoaniline materials can be used. Cameron further mention that such conduction materials can form flexible conductive films (page 22).

While a MW 2,000 is not Oka's preferred range, the eventual loss of flexibility is not viewed as a factor to prevent one of ordinary skill in the art from using lower MW materials when the teaching of Cameron clearly teaches that lower MW materials which overlap applicants' range are also effective flexible conductive films.

Applicant argues that there is no motivation to combine Oka and Nalwa with Kin.

The examiner counters that Oka and Nalwa teach the claimed materials as charge transporting materials. Kin clearly shows that the same class of material is commonly used in electroluminescent devices. As Kin discloses a similar type of conducting polyaniline material, it would have been obvious to a person of ordinary skill in the art at the time of the invention to have used the polyaniline materials of Oka and Nalwa as a thin film layer in an electroluminescent device, absent unexpected results.

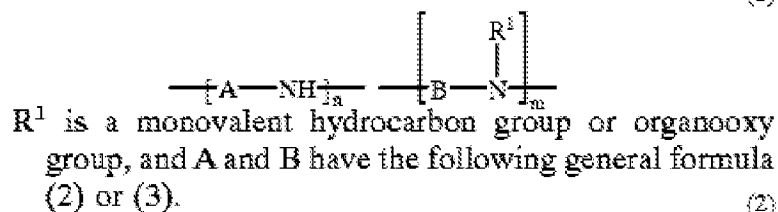
Conclusion

The prior art made of record and not relied upon is considered pertinent to applicant's disclosure. Kido (US 6,632,544) discloses charge transporting varnish represented by Formula K-1:

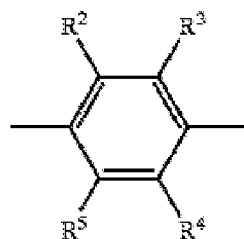
Art Unit: 1786

An aromatic amine derivative of the following general formula (1) is useful to form an auxiliary carrier transporting layer in an electroluminescent device.

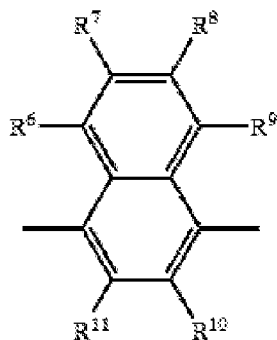
(1)



(2)



(3)



The MW of Formula K-1 is number average molecular weight is 250 to 100,000.

Any inquiry concerning this communication or earlier communications from the examiner should be directed to GREGORY CLARK whose telephone number is (571)270-7087. The examiner can normally be reached on M-Th 7:00 AM to 5 PM Alternating Fri 7:30 AM to 4 PM and Off.

If attempts to reach the examiner by telephone are unsuccessful, the examiner's supervisor, Larry Tarazano can be reached on (571) 272-1515. The fax phone number for the organization where this application or proceeding is assigned is 571-273-8300.

Art Unit: 1786

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/D. Lawrence Tarazano/
Supervisory Patent Examiner, Art Unit 1786

GREGORY CLARK/GDC/
Examiner
Art Unit 1786